

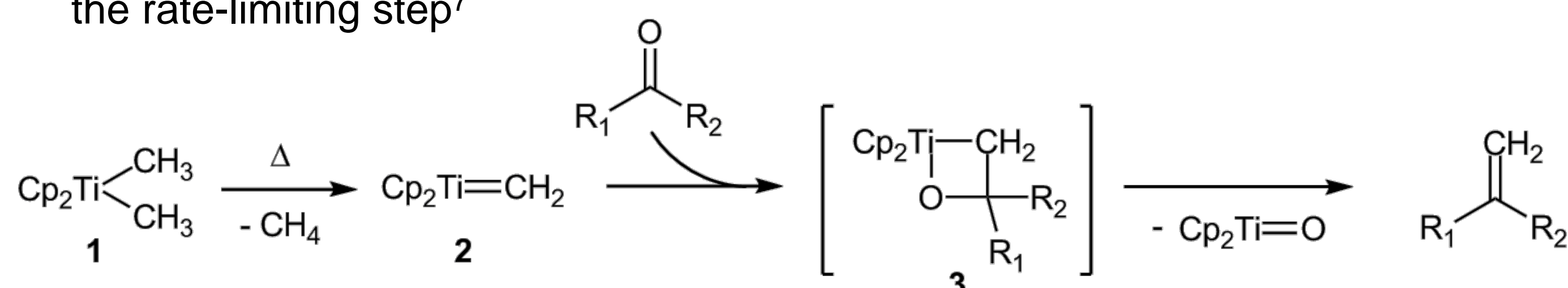
Petasis Olefination in a Continuous Flow Microwave Reactor: *exo*-Glycols from Sugar-Lactones

Introduction

Owing to their distinct nucleophilic reactivity, enol ethers and, in particular, *exo*-glycols are considered interesting synthons for further conversion, as for example to spiroketals¹ or C-glycosides². Commonly applied olefination reactions, e.g. Wittig-, Horner-, Julia- or Peterson-olefination³, require basic reaction conditions not applicable to the conversion of carboxylic esters to enol ethers. The Tebbe olefination⁴ is useful for the synthesis of enol ethers from esters. However, the required reagent is very sensitive to moisture. In contrast, the Petasis olefination⁵ of esters not only provides non-basic conditions, the reagent dimethyltitanocene **1** dissolved in toluene/THF also is stable towards moisture and air. It can be stored at + 4 °C over a longer period without decomposition⁶. Upon heating the molecule eliminates methane and forms the active titan carbenoid species **2**, which can react with the carbonyl compound in a cycloaddition reaction to form the more or less stable titanacycle **3**. After cycloreversion the olefinated product is released.

Reaction under batch conditions

- Substrate and reagent are dissolved in toluene/THF and heated in the dark under inert atmosphere at 65-80 °C for several hours
- Methane-elimination to form the reactive species starts at about 65 °C and is the rate-limiting step⁷



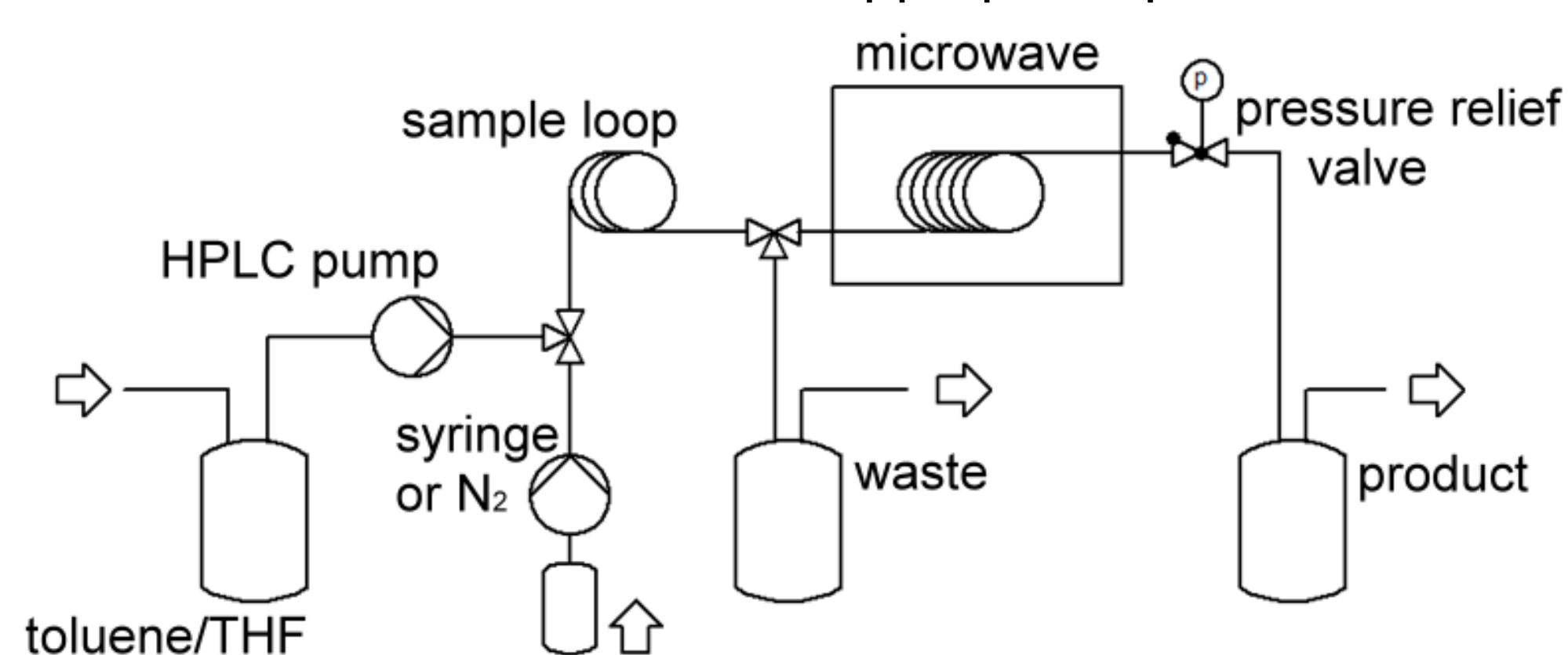
Increase of the temperature should strongly enhance the reaction rate limited by the boiling point of THF of 66 °C under atmospheric pressure.

Continuous flow microwave reactor under pressure (8 bar)

→ Optimization of crucial parameters with small amount of chemicals possible

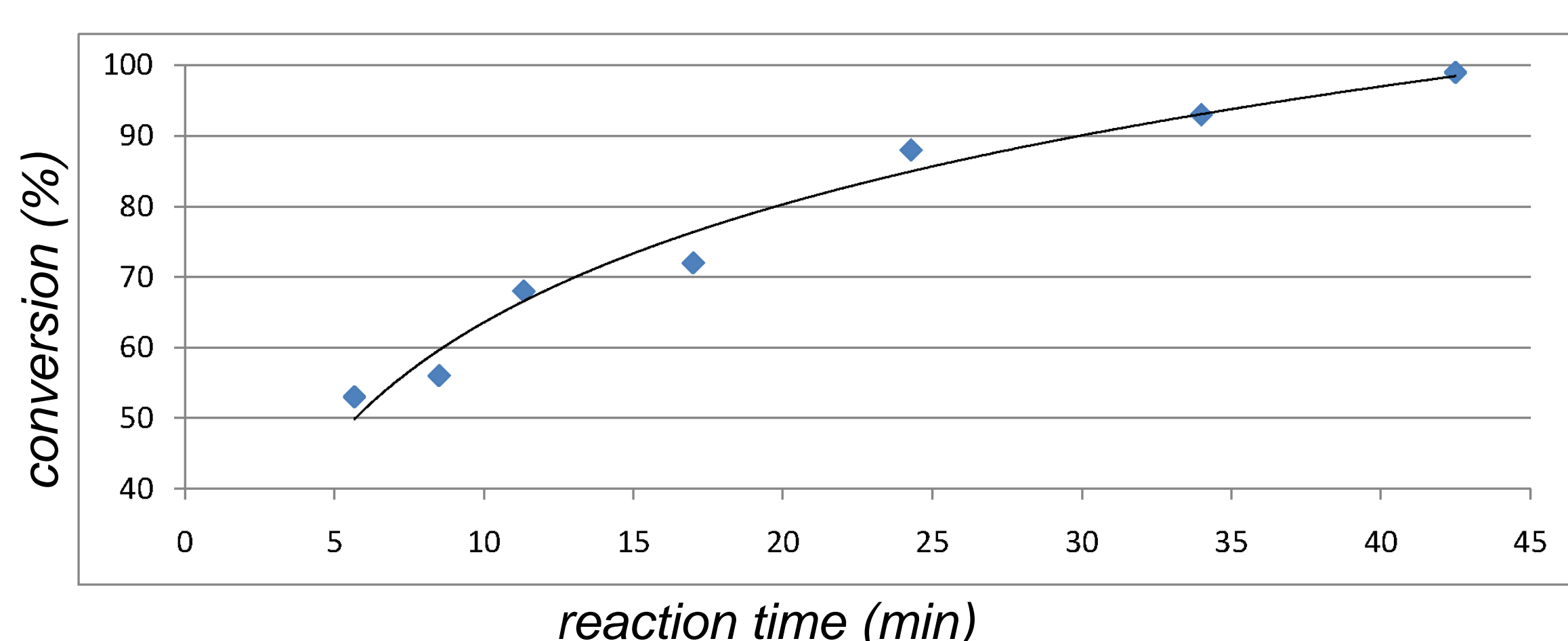
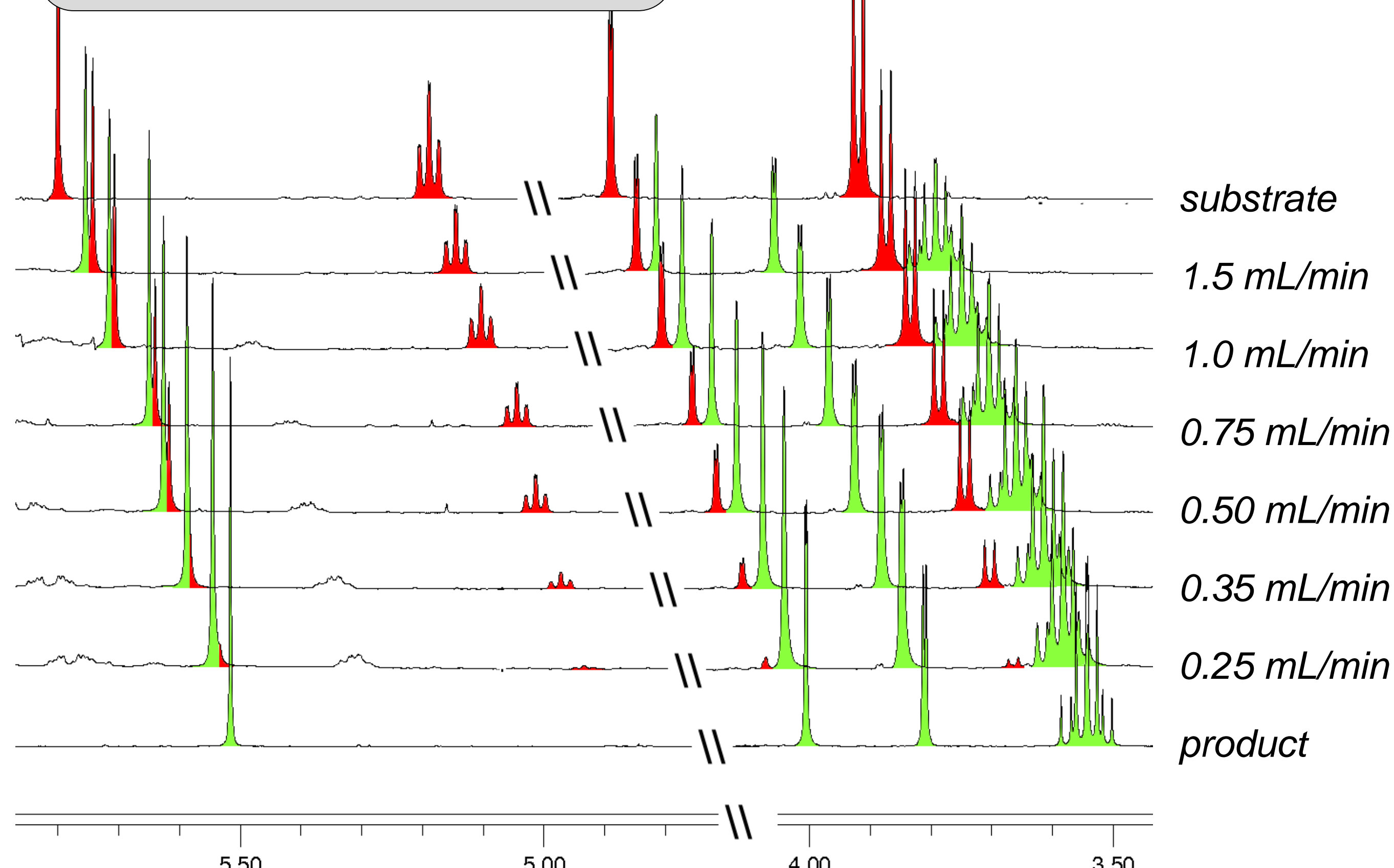
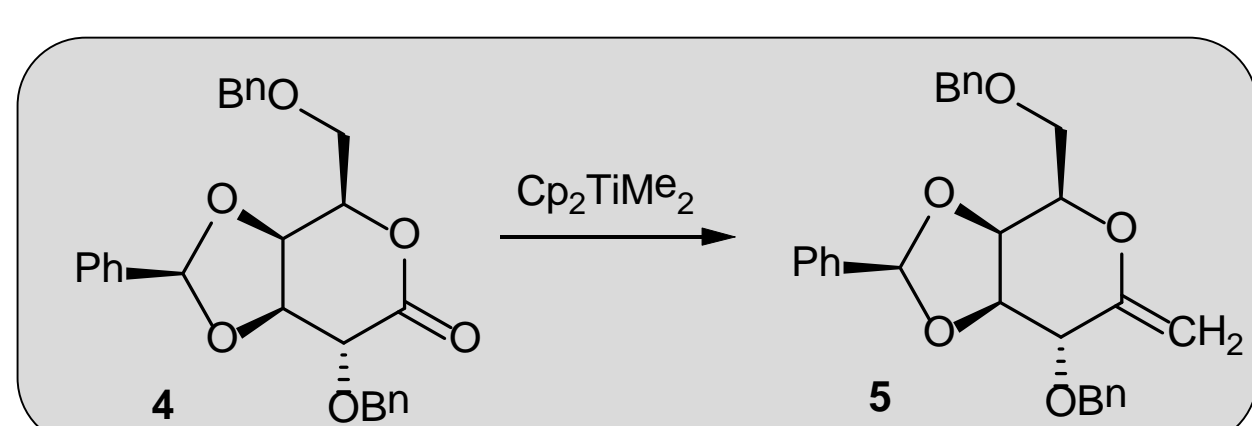
Continuous flow setup

- HPLC pump for delivering solvents and reactants
- Sample loop: capillaries (PTFE, 3.2 mm OD, 1.6 mm ID), total volume 6 mL
- Flow reactor (total volume 8.5 mL)
- Microwave oven for heating the reaction mixture
- Pressure relief valve to ensure the appropriate pressure



Reaction under continuous flow conditions

- The olefination reaction can be monitored by NMR after each run as illustrated for the galactose-derived sugar-lactone **4**



Preparative application

- Glucose-derived building block **6** reacts remarkably faster under similar conditions → complete conversion already reached after 5.67 min (at 1.5 mL/min)
- The formation of **2** by α -elimination must be completed very fast in both cases
- Contrary to the assumptions under batch conditions, the rate determining step must be the bimolecular reaction of **2** with the substrate
- Higher concentrations = faster reaction → remarkably accelerated conversion by increasing the Petasis reagent concentration from 0.18 M to 0.46 M
- Lowering the reaction time in the best cases down to only 2.8 min

Reaction	Continuous Flow		Microwave Batch	
	Yield	Time	Yield	Time
D ⁻ galactose (8 steps) → 4 → 5	71 %	5.7 min	70 %	30 min
D ⁻ glucose (6 steps) → 6 → 7	65 %	5.7 min	67 %	30 min
D ⁻ galactose (6 steps) → 8 → 9	52 %	5.7 min	50 %	30 min
D ⁻ galactose (8 steps) → 10 → 11	70 %	5.7 min	70 %	30 min
D ⁻ galactose (7 steps) → 12 → 13	50 %	5.7 min	51 %	30 min
L ⁻ glucose (6 steps) → 14 → 15	74 %	8.5 min	74 %	30 min
Thermal batch:			74 %	720 min ⁸

- Preparative application of the continuous flow reaction even on gram-scale
- Scale-up by simply increasing the operation time of the reactor
- No changes regarding conversion or yield

Summary

The olefination of complex sugar-derived lactones to *exo*-glycols was achieved under continuous flow conditions. This procedure allowed an optimization of the conversion by regulation of the residence time and the concentration on milligram-scale, before the ready scale-up of the reaction to a gram-scale production. It can be concluded from these results, that micro reactors are also efficient in preparative microwave-assisted conversions of other sensitive carbonyl substrates under continuous flow conditions.

References

- Brand C.; Rauch, G.; Zanoni, M.; Dittrich, B.; Werz, D. B. *J. Org. Chem.*, **2009**, *74*, 8779 and literature cited therein.
- Woodward, H.; Smith, N.; Gallagher, T. *Synlett*, **2010**, 869.
- (a) Maryanoff, B. E.; Reitz, A. B. *Chem. Rev.* **1989**, *89*, 863-927; (b) Mikolajczyk, M.; Balczewski, P. *New Aspects in Phosphorous Chemistry II*; Springer Berlin, Heidelberg, **2003**; 161-214; (c) Blakemore, P. R. *J. Chem. Soc., Perkin Trans. 1*, **2002**, 2563-2585; (d) van Staden, L. F.; Gravestock, D.; Ager, D. *J. Chem. Soc. Rev.*, **2002**, *31*, 195-200.
- Tebbe, F. N.; Parshall, G. W.; Reddy, G. S. *J. Am. Chem. Soc.*, **1978**, *100*, 3611.
- Petasis, N. A.; Bzowej, E. I. *J. Am. Chem. Soc.* **1990**, *112*, 6392; (b) Petasis, N. A.; Patane, M. A. *Tetrahedron Lett.* **1990**, *31*, 6799; (c) Petasis, N. A.; Lu, S.-P. *Tetrahedron Lett.* **1996**, *37*, 141; (d) Csuk, R.; Glänzer, B. I. *Tetrahedron* **1991**, *47*, 1655; (e) Hartley, R. C.; McKiernan, G. J. *J. Chem. Soc., Perkin Trans. 1*, **2002**, 2763-2793; (f) Hartley, R. C.; Ji, J. F.; Main, C. A.; McKiernan, G. J. *Tetrahedron*, **2007**, *63*, 4825-4864.
- (a) Payack, J. F.; Hughes, D. L.; Cai, D. W.; Cottrell, I. F.; Verhoeven, T. R. *Org. Prep. Proc. Int.*, **1995**, *27*, 707; (b) Payack, J. F.; Huffman, M. A.; Cai, D. W.; Hughes, D. L.; Collins, P. C.; Johnson, B. K.; Cottrell, I. F.; Tuma, L. D. *Org. Process Res. Dev.*, **2004**, *8*, 256.
- Hughes, D. L.; Payack, J. F.; Cai, D. W.; Verhoeven, T. R.; Reider, P. J. *Organometallics*, **1996**, *15*, 663; (b) Meurer, E. C.; Santos, L. S.; Pilli, R. A.; Eberlin, M. N. *Org. Lett.* **2003**, *5*, 1391-1394.
- Gallagher, T.; Cook, M. J.; Declan, W. F. *Tetrahedron Lett.* **2005**, *46*, 297-300